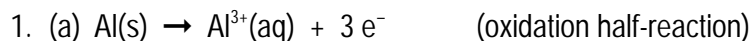
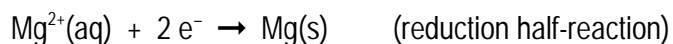
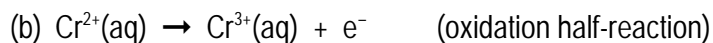


ELECTROCHEMISTRY REVIEW AND PRACTICE 1

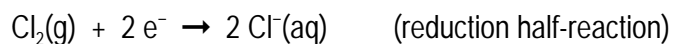
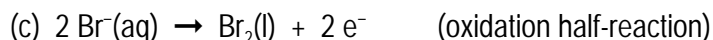
ANSWERS



The reaction is spontaneous [Al(s) is the reducing agent; $\text{Pb}^{2+}(\text{aq})$ is the oxidizing agent].



The reaction is not spontaneous [$\text{Cr}^{2+}(\text{aq})$ is the reducing agent; $\text{Mg}^{2+}(\text{aq})$ is the oxidizing agent].



The reaction is spontaneous [$\text{Br}^-(\text{aq})$ is the reducing agent; $\text{Cl}_2(\text{g})$ is the oxidizing agent].

